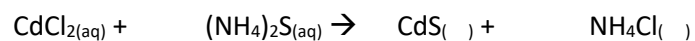
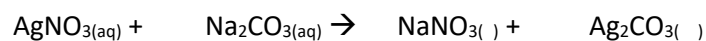


Questions:

1. Balance and predict the solubility of each product for the following two reactions:



2. Complete the following reactions, balance, and predict the states of the products:



3. What mass of solid aluminum hydroxide can be produced when 50.0 mL of 0.200 M of aluminum nitrate is added to 200.0 mL of 0.100 M of potassium hydroxide?
What is the concentration of aluminum ions left over in solution after the reaction is complete?
4. If 10.0 mL of a 0.30 mol/L sodium phosphate solution reacts with 20.0 mL of a 0.20 mol/L lead(II) nitrate solution, what is the mass of the precipitate?
5. How many grams of precipitate are formed with 100 mL of 0.100M potassium sulfate and 100 mL of 0.200M barium nitrate? Note – it's a double displacement reaction.

Write out the complete and net ionic equations after the calculations.